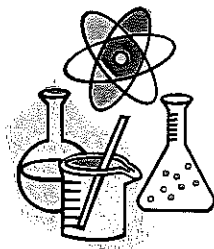


Key

Name _____
Date _____
Period _____



Chemistry Study Guide

1. Complete the Chart below and answer questions a-d:

Subatomic Particle	Charge	Relative Mass	Location in atom
Proton	+	1 amu	nucleus
Neutron	none	1 amu	nucleus
electron	-	0	electron cloud

a. Which subatomic particle is the atomic number on the Periodic Table?

proton

b. Which two subatomic particles tell us the atomic mass of an element?

proton + neutrons

c. Atoms with the same number of protons but different number of neutrons are called isotopes.

d. In a neutral atom, the number of protons is equal to the number of electrons.

2. Use the element to answer the following 2 questions:

34
Se
78.96

a. How many protons does Selenium have? 34

b. How many neutrons does Selenium have? 45

$$79 - 34 =$$

Groups	different	Clouds
Periods	Electrons	Mass
viscosity	Similar	Isotopes
electrons	Valence electrons	Ions

Use the Word Bank above and answer questions 3-8.

3. A _____ is a change in size, shape, or state.
4. Elements of the same group or family have similar properties.
5. The resistance of a fluid's flow is known as viscosity.
6. Negatively charged particles called electrons orbit the nucleus.
7. The most modern model of atoms show clouds of electrons around the nucleus.
8. Isotopes are atoms with the same number of protons but different number of neutrons.
9. Atoms that are electrically charged are considered to be ions.
10. Periods on the Periodic table run in rows.
11. Groups on the Periodic Table run as columns.
12. Valence Electrons can give you hints on the reactivity and chemical properties of a substance.